Table IIIc. Partial quantities for Mg in the liquid phase at 1400 K.

x_{Mg}	ΔG_{Mg} [J/mol]	ΔH_{Mg} [J/mol]	ΔS_{Mg} [J/(mol·K)]	$G^{\mathrm{E}}_{\mathrm{Mg}}$ [J/mol]	$S^{\mathrm{E}}_{\mathrm{Mg}}$ [J/(mol·K)]	a_{Mg}	γ_{Mg}
0.000	$-\infty$	−45145	∞	−38503	−4.744	0.000	0.037
0.100	−55340	−33916	15.302	−28537	−3.843	0.009	0.086
0.200	−39187	−24704	10.346	−20453	−3.036	0.035	0.173
0.300	−28070	−17310	7.686	−14056	−2.325	0.090	0.299
0.400	−19814	−11539	5.911	−9148	−1.708	0.182	0.456
0.500	−13603	−7195	4.577	−5535	−1.186	0.311	0.622
0.600	−8965	−4081	3.488	−3019	−0.759	0.463	0.772
0.700	−5555	−2001	2.539	−1403	−0.427	0.620	0.886
0.800	−3090	−758	1.666	−493	−0.190	0.767	0.959
0.900	−1317	−157	0.829	−90	−0.047	0.893	0.992
1.000	0	0	0.000	0	0.000	1.000	1.000

Reference state: Mg(liquid)

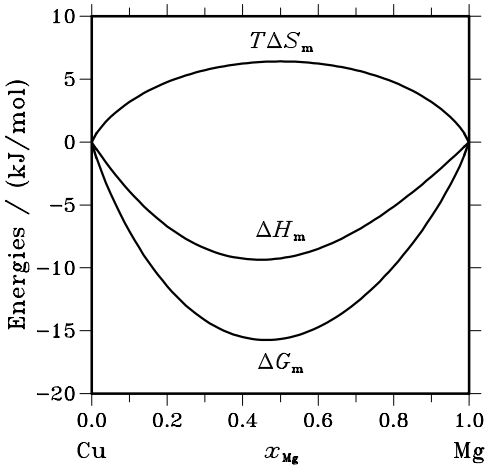

Fig. 2. Integral quantities of the liquid phase at T=1400 K.

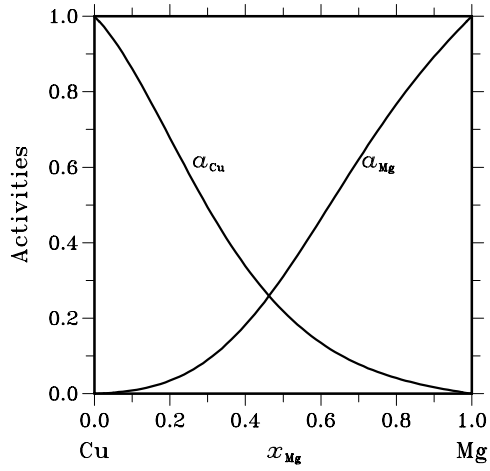

Fig. 3. Activities in the liquid phase at T=1400 K.

Table IV. Standard reaction quantities at 298.15 K for the compounds per mole of atoms.

Compound	x_{Mg}	$\Delta_{\mathrm{f}}G^{\circ}$ / (J/mol)	$\Delta_{\mathrm{f}}H^{\circ}$ / (J/mol)	$\Delta_{\mathrm{f}}S^{\circ}$ / (J/(mol·K))	$\Delta_{\mathrm{f}}C^{\circ}_{P}$ / (J/(mol·K))
$C15$	0.333	−11278	−10908	1.241	−1.432
Cu_1Mg_2	0.667	−9355	−9540	−0.622	0.000

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